HW08 - Kinetics

Question 1 0.5 pts

Consider the reaction:

$2O_3(g) \longrightarrow 3O_2(g)$ rate = $k[O_3]^2[O_2]^{-1}$

What is the overall order of the reaction and the order with respect to $[O_3]$?

- 2 and 2
- -1 and 3
- 3 and 2
- 1 and 2

Question 2 0.5 pts

When the reaction below:

$3NO(g) \longrightarrow N_2O(g) + NO_2(g)$

is proceeding under conditions such that 0.015 mol/L of N_2O is being formed each second, the rate of the overall reaction is _____ and the rate of change for NO is _____.

- 0.015 M/s; -0.005 M/s
- 0.030 M/s; -0.005 M/s
- 0.015 M/s; -0.045 M/s
- 0.015 M/s; 0.045 M/s

Question 3 1 pts

What is the rate law for the reaction below:

$A + B + C \longrightarrow D$

if the following data were collected?

Exp	$[A]_0$	$[B]_0$	$[C]_0$	Initial Rate
1	0.4	1.2	0.7	2.32×10^{-3}
2	1.3	1.2	0.9	7.54×10^{-3}
3	0.4	4.1	0.8	9.25×10^{-2}
4	1.3	1.2	0.2	7.54×10^{-3}

- rate = 3.36×10^{-3} $[A]$ $[B]^3$
- rate = 1.49×10^{-3} $[B]^3$ $[C]$
- rate = 5.37×10^{-3} $[A]$ $[B]^3$
- rate = 1.79×10^{-3} $[B]^2$ $[C]$
- rate = 4.48×10^{-3} $[A]$ $[B]^2$ $[C]$

Question 4 1 pts

A chemical reaction is expressed by the balanced chemical equation:

$A + 2B \longrightarrow C$

Consider the data below:

exp	$[A]_0$	$[B]_0$	initial rate (M/min)
1	0.15	0.15	0.00110363
2	0.15	0.3	0.0044145
3	0.3	0.3	0.008829

Find the rate law for the reaction.

- rate = k $[B]^2$
- rate = k $[A]^2$ $[B]$
- rate = k $[A]$ $[B]^2$
- rate = k $[A]$ $[B]$

Question 5 0.5 pts

Calculate the value of the rate constant (k) for the reaction in question 4.

- 0.00736
- 0.00110
- 0.000166
- 0.327

Question 6 0.5 pts

If the initial concentrations of both A and B are 0.31 M for the reaction in questions 4 and 5, at what initial rate is C formed?

- 0.00974 M/min
- -0.00974 M/min
- 0.0314 M/min
- 0.101 M/min

Question 7 — 1 pts

We know that the rate expression for the reaction below:

$2NO + O_2 \longrightarrow 2NO_2$

at a certain temperature is rate = $[NO]^2 [O_2]$. We carry out two experiments involving this reaction at the same temperature, but in the second experiment the initial concentration of NO is doubled while the initial concentration of O_2 is halved. The initial rate in the second experiment will be how many times that of the first?

- 2
- 8
- 4
- 1

Question 8 — 1 pts

Consider the data collected for a chemical reaction between compounds A and B that is first order in A and first order in B:

rxn	$[A]_0$	$[B]_0$	rate (M/s)
1	0.2	0.05	0.1
2	?	0.05	0.4
3	0.4	?	0.8

From the information above for 3 experiments, determine the missing concentrations of A and B. Answers should be in the order [A] then [B].

- 0.40 M; 0.10 M
- 0.20 M; 0.80 M
- 0.80 M; 0.10 M
- 0.80 M; 0.20 M
- 0.40 M; 0.20 M

Question 9 — 0.5 pts

For a reaction that is zero-order overall...

- the rate does not change during the reaction.
- the reactant concentration does not change with time.
- the rate constant is zero.
- the activation energy is zero.

Question 10 — 0.5 pts

Consider the reaction below:

$A + B \longrightarrow C$

If it is 1st order in A and 0th order in B, a plot of ln[A] vs time will have a slope that is...

- slowly increasing.
- decreasing exponentially.
- constant.
- increasing exponentially.

Question 11 — 1 pts

Consider the reaction below:

$H_2CO_3(aq) \longrightarrow CO_2(aq) + H_2O(l)$

If it has a half-life of 1.6 sec, how long will it take a system with $[H_2CO_3]_0$ of 2M to reach $[H_2CO_3]$ of 125mM?

- Not enough information is given.
- 3.2 sec
- 6.4 sec
- 2.9 sec

Question 12 — 1 pts

At a certain fixed temperature, the reaction below:

$A(g) + 2B(g) \longrightarrow AB_2(g)$

is found to be first order in the concentration of A and zeroth order in the concentration of B. The reaction rate constant is $0.05s^{-1}$. If 2.00 moles of A and 4.00 moles of B are placed in a 1.00 liter container, how many seconds will elapse before the concentration of A has fallen to 0.30 moles/liter?

- 10.22 sec
- 37.94 sec
- There is not enough information to answer.
- 2.83 sec

Question 13 — 0.5 pts

The reaction below:

$A \longrightarrow$ products

is observed to obey first-order kinetics. Which of the following plots should give a straight line?

- ln[A] vs t
- ln[A] vs k
- [A] vs t
- [A] vs t^{-1}
- ln[A] vs t^{-1}
- [A] vs k
- ln[A] vs k^{-1}

Question 14 1 pts

For the reaction below:

cyclobutane(g) $\longrightarrow$ 2ethylene(g)

at 800K, a plot of ln[cyclobutane] vs t gives a straight line with a slope of -1.6 s^{-1}. Calculate the time needed for the concentration of cyclobutane to fall to 1/16 of its initial value.

- 0.63 sec
- 1.3 sec
- 1.7 sec
- 1.6 sec

Question 15 1 pts

The initial concentration of the reactant A in a first-order reaction is 1.2 M. After 69.3 sec, the concentration has fallen to 0.3 M. What is the rate constant k?

- 0.02 s^{-1}
- 0.2 s^{-1}
- 0.01 s^{-1}
- not enough information

Question 16 1 pts

A reaction is found to be first order with respect to one of the reactant species, A. When might a plot of ln[A] vs time NOT yield a straight line?

- if the reaction comes to equilibrium
- if the reaction has any significant backward rate
- when the rate also depends on the concentration of another reactant as well
- All of the other answers could be correct.

Question 17 0.5 pts

Consider the following elementary reactions:

a) NO + O_3 $\longrightarrow$ NO_2 + O_2

b) CS_2 $\longrightarrow$ CS + S

c) O + O_2 + N_2 $\longrightarrow$ O_3 + N_2

Identify the molecularity of each reaction respectively.

- bimolecular, unimolecular, termolecular
- all three elementary reactions are bimolecular
- tetramolecular, termolecular, pentamolecular
- it is impossible to know without knowing the overall reaction for each

Question 18 1 pts

A and B react to form C according to the single step reaction below:

A + 2B $\longrightarrow$ C

Which of the following is the correct rate equation for [B] and the correct units for the rate constant of this reaction?

- $\frac{\Delta[B]}{\Delta t} = -\frac{2k[A][B]}{[C]}$; $\frac{1}{M \cdot s}$
- $\frac{\Delta[B]}{\Delta t} = -2k[A][B]$; $\frac{1}{M \cdot s}$
- $\frac{\Delta[B]}{\Delta t} = -2k[A][B]^2$; $\frac{1}{M^2 \cdot s}$
- $\frac{\Delta[B]}{\Delta t} = -k[A][B]^2$; $\frac{1}{M^2}$

Question 19 1 pts

Consider the mechanism below:

NO_2 + F_2 $\longrightarrow$ NO_2F + F k_1, slow

F + NO_2 $\longrightarrow$ NO_2F k_2, fast

What is the rate law?

- rate = $k_1[NO_2][F_2]$
- rate = $k_2[NO_2][F]$
- rate = $k_1[NO_2F][F_2]$
- rate = $k_1k_2[NO_2]^2$
- rate = $k_2[NO_2]^2$

Question 20 1 pts

Determine the overall balanced equation for a reaction having the following proposed mechanism:

Step 1: B_2 + B_2 $\longrightarrow$ E_3 + D slow

Step 2: E_3 + A $\longrightarrow$ B_2 + C_2 fast

and write an acceptable rate law.

- A + B_2 $\longrightarrow$ C_2 + D; rate = $k[B_2]^2$
- A + B_2 $\longrightarrow$ C_2 + D; rate = $k[A][B_2]$
- 2B_2 $\longrightarrow$ E_3 + D; rate = $k[B_2]^2$
- E_3 + A $\longrightarrow$ B_2 + C_2; rate = $k[E_3][A]$

Question 21 1 pts

Consider the reaction below:

$H_2(g) + I_2(g) \longrightarrow 2HI(g)$

The proposed mechanism of this reaction is:

$I_2 \rightleftharpoons 2I$ k_1, k_{-1}(reverse rxn), fast

$2I + H_2 \longrightarrow 2HI$ k_2, slow

What is the rate of the overall reaction?

- $rate = k_1 k_2 [I_2][H_2]$
- $rate = \frac{k_1 k_2}{k_{-1}}[I]^2[H_2]$
- $rate = k_2[I]^2[H_2]$
- $rate = \frac{k_{-1} k_2}{k_1}[I_2][H_2]$
- $rate = \frac{k_1 k_2}{k_{-1}}[I_2][H_2]$

Question 22 1 pts

A reaction rate increases by a factor of 655 in the presence of a catalyst at 37°C. The activation energy of the original pathway is 106 kJ/mol. What is the activation energy of the new pathway, all other factors being equal?

- 16,600 J/mol
- 89.3 kJ/mol
- 89.3 J/mol
- 16,600 kJ/mol

Question 23 1 pts

A given reaction has an activation energy of 24.52 kJ/mol. At 25°C, the half-life is 4 minutes. At what temperature will the half-life be reduced to 20 seconds?

- 100°C
- 125°C
- 115°C
- 150°C

Question 24 1 pts

For the reaction below:

$HO(g) + H_2(g) \longrightarrow H_2O(g) + H(g)$

a plot of lnK vs 1/T gives a straight line with a slope equal to -5.1×10^3 K. What is the activation energy for this reaction?

- 12 kJ/mol
- 42 kJ/mol
- 5.1 kJ/mol
- 98 kJ/mol

Question 25 1 pts

A certain reaction has an activation energy of 0.8314 kJ/mol and a rate constant of 2.718 s^{-1} at -73°C. At -173°C, which expression for the rate constant is correct?

- $\ln(k_2) = 0.5$
- $\ln(k_2) = -0.5$
- $\ln(k_2) = 1$
- $\ln(k_2) = 1.5$

Question 26 1 pts

A food substance kept at 0°C becomes rotten (as determined by a good quantitative test) in 8.3 days. The same food rots in 10.6 hours at 30°C. Assuming the kinetics of the microorganisms enzymatic action is responsible for the rate of decay, what is the activation energy for the decomposition process? Hint: Rate varies INVERSELY with time; a faster rate produces a shorter decomposition time.

- 23.4 kJ/mol
- 67.2 kJ/mol
- 2.34 kJ/mol
- 0.45 kJ/mol

Question 27 1 pts

A catalyst...

- speeds up the reaction but does not change K.
- increases K to favor product formation.
- changes the reaction mechanism to ensure that K is increased.
- speeds up the reaction and increases K to favor product formation.

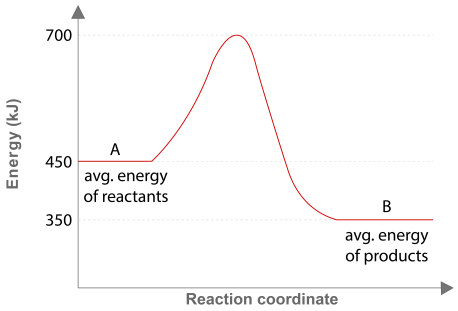
Question 28 1 pts

All else being equal, a reaction with a higher activation energy compared to one with a lower activation energy will...

- proceed faster.
- be more endothermic.
- proceed slower.
- be more exothermic.

Question 29 — 1 pts

Consider the potential energy diagram below:



What is the change in enthalpy (ΔH) for the reaction A ⟶ B?

- -100 kJ
- 350 kJ
- 100 kJ
- -350 kJ

Question 30 — 1 pts

Consider the potential energy diagram in question 39. What is the activation energy (E_a) for the reaction?

- 350 kJ
- 100 kJ
- 200 kJ
- 250 kJ

Question 31 — 1 pts

Which of the following statements is TRUE?

- If the exponents in the rate-law do not match the coefficients in the balanced chemical equation, then we know that the reaction takes place in one step.
- The exponents in the rate-law must match the coefficients in the balanced chemical equation for the reaction.
- If the exponents in the rate-law do not match the coefficients in the balanced equation, then we know that the reaction does not take place in one step.
- The rate-law for a reaction can be predicted from the balanced chemical equation.

Question 32 — 1 pts

"Reaction mechanisms usually involve only unimolecular or bimolecular steps."

Is this statement true or false?

- True, because collisions of higher molecularity would occur too infrequently to account for an observed rate.
- True, because steps of higher molecularity would not be compatible with observed reaction rate laws.
- False.
- True, because the activation energy for collisions of higher molecularity would be too great.

Question 33 — 1 pts

Which of the following is/are ALWAYS true concerning collision and transition state theory?

I) Transition states are short-lived.

II) A balanced reaction shows which species must collide for the reaction to occur.

III) Intermediates are short-lived.

- I and III
- All are true.
- II and III
- I only
- III only
- II only

Question 34 — 1 pts

Consider the following reaction mechanism:

1) $Cl_2 + Pt \longrightarrow 2Cl + Pt$

2) $Cl + CO + Pt \longrightarrow ClCO + Pt$

3) $Cl + ClCO \longrightarrow Cl_2CO$

Overall: $Cl_2 + CO \longrightarrow Cl_2CO$

Which species is/are intermediates?

- Pt, Cl, ClCO
- Pt
- ClCO
- Pt, Cl
- Cl, ClCO